2.2.1 The institution assesses the learning levels of the students and organizes special Programmes for advanced learners and slow learners

Special activities for slow learners

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Circular

This is to inform all First Year B Pharm students that Remedial classes will be conducted from 29/11/2019.

Find the below attached list of students with the mentor details for remedial classes.

Krupanidhi College of Pharmacy IQAC Date: 18/10/2022 Bangalore-35

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KRUPANIDHI COLLEGE OF PHARMACY, BANGALORE

Timetable for Remedial Class of I Sem B.Pharm (2019-20)

Subject	Viva /Exam	Dates & time	Time	Mrs Sayani	Mrs Arnika	Dr. C. Sudha	Mrs Gargi
Pharmaceutics	VIVA/ TESTS	29/11/19	9 – 4.30pm	Group A	Group B	Group C	Group D
		30/11/19	9 – 4.30pm	Group A	Group B	Group C	Group D
		2/12/19	9 –1.30 pm	Group A	Group B	Group C	Group D
		3/12/19	9 -12 noon	Group A	Group B	Group C	Group D
	PREFINAL EXAM	3/12/2019 (1pm to 4 pm)		Additional Exam	03/12/2019 (2.30 pm to 4 pm)		
PharmInorganic Chemistry	VIVA/ TESTS	4/12/19	9 – 4.30pm	Group B	Group A	Group C	Group D
		5/12/19	9 – 4.30pm	Group B	Group A	Group C	Group D
		6/12/19	9 – 4.30pm	Group B	Group A	Group C	Group D
		7/12/19	9 -12 noon	Group B	Group A	Group C	Group D
	PREFINAL EXAM	9/12/2019(1pm to 4 pm)		Additional Exam	09/12/2019 (2.30 pm to 4 pm)		
Pharm analysis	VIVA/ TESTS	10/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
		11/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
		12/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
		13/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
	PREFINAL EXAM	14/12/2019(10 am – 1 pm)		Additional Exam	14/12/2019 (11.30am to 1 pm)		
HAP	VIVA/ TESTS	16/12/19	9 – 4.30pm	Group D	Group B	Group C	Group A
		17/11/19	9 – 4.30pm	Group D	Group B	Group C	Group A
		18/11/19	9 – 4.30pm	Group D	Group B	Group C	Group A

Krupanidhi College of Pharmacy IQAC Date: 18/10/2022 Bangalore-35

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Subject	Viva /Exam	Dates & time	Time	Mrs Sayani	Mrs Arnika	Dr. C. Sudha	Mrs Gargi
Pharmaceutics	VIVA/ TESTS	29/11/19	9 – 4.30pm	Group A	Group B	Group C	Group D
		30/11/19	9 – 4.30pm	Group A	Group B	Group C	Group D
		2/12/19	9 –1.30 pm	Group A	Group B	Group C	Group D
		3/12/19	9 -12 noon	Group A	Group B	Group C	Group D
	PREFINAL EXAM	**3/12/2019 (1pm to 4 pm)**		**Additional Exam**	**03/12/2019 (2.30 pm to 4 pm)**		
PharmInorganic Chemistry	VIVA/ TESTS	4/12/19	9 – 4.30pm	Group B	Group A	Group C	Group D
		5/12/19	9 – 4.30pm	Group B	Group A	Group C	Group D
		6/12/19	9 – 4.30pm	Group B	Group A	Group C	Group D
		7/12/19	9 -12 noon	Group B	Group A	Group C	Group D
	PREFINAL EXAM	**9/12/2019(1pm to 4 pm)**		**Additional Exam**	**09/12/2019 (2.30 pm to 4 pm)**		
Pharm analysis	VIVA/ TESTS	10/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
		11/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
		12/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
		13/12/19	9 – 4.30pm	Group D	Group B	Group A	Group C
	PREFINAL EXAM	**14/12/2019(10 am – 1 pm)**		**Additional Exam**	**14/12/2019 (11.30am to 1 pm)**		
HAP	VIVA/ TESTS	16/12/19	9 – 4.30pm	Group D	Group B	Group C	Group A
		17/11/19	9 – 4.30pm	Group D	Group B	Group C	Group A
		18/11/19	9 – 4.30pm	Group D	Group B	Group C	Group A

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KRUPANIDHI COLLEGE OF PHARMACY, BANGALORE

I SEM B.PHARM (2019-20)

REMEDIAL CLASS

BATCH LIST

Sl No	GROUP A	GROUP B	GROUP C	GROUP D
	GARGI	ARNIKA	SUDHA MAM	SAYANI MAM
1.	Abhishek	Cebin	Aleena	Brahmhotri
2.	Usman	Deep	Shylashree	Deepa
3.	Ajith	Shivsekhar	Bikando	Gururaj
4.	Shreyas	Divya M	Chandana	Hamsa
5.	Vikram	Ganesh	Durga Prasad	Kavya S
6.	Rohul	Balu Prashanth	Charitha	Daphisa
7.	Ankit	Vikas	Divya P	Keerthana Iyer
8.	Christy	Manisha	Angela	Shubha
9.	Sanjay	Kavyashree	Vishal	Shalini N
10.	Rounak	Pratistha	Jasmitha	Nochutenu
11.	Siravana	Dhruva	Ramya M	Vasanth
12.	Sumesh	Likitha	Priyanka s	Pradeep
13.	Hrithik kumar	Samreen	Lavanya	Supreetha
14.	Divakar	Rachna	Nikhil Gowda	Chandu
15.	Pranjal	Pratibha	Ramya N	Spoorthi
16.	Mahesh	Spandita	Lalramhua	Vedanayagam

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17.	Vigneswaram	Nirma	Roja	Talisunep
18.	Kunal	Poonam	Madhumita	Gagan
19.	Kikon	Shalini P	Sahana Priya	Baby Nandini
20.	Mohit	Lekha	Sonam	Suresh
21.	Nitesh	Rakshitha	Pratishtha Phukan	Banlumlang
22.	Pavan			
23.	Lakshmish			
24.	Prasanna			
25.	Prashanth			
26.	C lalthianglima			
27.	Prakash			
28.	sudeep			
29.	Stephen Ashwin			
30.	Karthick			
31.	Pooja			
32.	Syed Jommy			
33.	Illakkiya			
34.	Nimisha T			

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KRUPANIDHI COLLEGE OF PHARMACY, BANGALORE
I Sessional Exam of I SEM B.Pharm 2019-20
Sub: PHARMACEUTICAL ANALYSIS

Max. Marks: 30

Long Essay (Answer any One) **(1x10=10)**

1. What are errors and classify them (5M) and describe how you minimize the errors.(5M)
2. Classify Acid – Base titrations and explain about different theories of indicators.

Short Essay (Answer any Two) **(02x05=10)**

3. What type of compounds are considered as primary standards and why? Give suitable examples.
4. A) Define accuracy, precision and significant figures.
   B) What are neutralization curves?
5. How do you prepare and standardize 0.1N Acetous perchloric acid.

Short notes (Answer all) **(05x02=10)**

6. Define the terms Normality and Molarity.
7. Define qualitative and quantitative analysis.
8. Name the solvents used in the non aqueous titrations.
9. Give the use of the following in non aqueous titrimetry
   a) Sodium methoxide
   b) Acetic acid
   c) Acetic anhydride
   d) Quinaldein red
10. How do you prepare 250ml 0.5M sodium thio sulfate.

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KRUPANIDHI COLLEGE OF PHARMACY, BANGALORE
II Sessional Exam of I SEM B.Pharm 2019-20
Sub: PHARMACEUTICAL ANALYSIS

Max. Marks: 30

Long Essay (Answer any One) **(1x10=10)**

1. Define Redox titrations. Give the examples for the Oxidising and Reducing agents and explain about the types of redox titrations.
2. Define complexometry and classify the complexing agents with examples and give the applications of complexometry.

Short Essay (Answer any Two) **(02x05=10)**

3. Explain about the principle, advantages and disadvantages of Volhard's and Modified Volhard's method.
4. Define polarography. Explain about the plorographic curve how it is plotted and mention the different areas.
5. Explain the principle, working, advantages and disadvantages of Dropping mercury electrode.

Short notes (Answer all)
(05x02=10)

6. What is standard hydrogen electrode?
7. What are masking and demasking agents?
8. What is the difference between chelates and complexes?
9. What is the difference between conductometry and potentiometry?
10. What are the conditions required in the process of digestion in gravimetry?

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Remedial classes portions for I SEM I B. Pharm (2019-20)

Subject: Pharmaceutical analysis

SESSION 1:

10/12/19 please ask the students write test by 3.00PM

10Marks

1. Define and classify errors? Describe the various methods to minimize the errors.
2. What are acid base titration? Explain the Neutralisation curve in acid base titration.
3. What are primary and secondary standards? Give examples of primary standards used in different types of titrations. Enlist the ideal properties of primary standard.
4. Explain the procedure for selection of indicators in the titration between strong acid and strong base using neutralisation curves.
5. Classify acid base titrations. Explain the Quinonoid theory of indicators with example.
6. What are Neutralization curves? Explain the selection of indicators in the titration between weak acid with strong base using neutralization curves.

5Marks

1. How do you calculate the equivalent weight and molecular weight of a substance? Give examples.
2. Briefly explain the different theories of indicators.
3. How do you prepare and standardise the following compounds a) 500ml of 0.1N hydrochloric acid b) 250ml of 0.1N sodium hydroxide. C) 500ml 0.1N Potassium permanganate d) 250ml 0.1N ceric ammonium sulfate
4. How do you prepare and standardise 0.1N perchloric acid solution?
5. What is leveling effect in non aqueous titrimetry? Explain in detail.
6. What is non-aqueous titration? Give the principle and procedure involved in estimation of Ephedrine Hydrochloride and sodium benzoate.
7. What is standardization? What type of substances should be standardised? How do you prepare 200ml of 0.5 N Oxalic acid solution?
8. What is pharmaceutical analysis? Explain different types of analysis. What is its scope in pharmacy?

2Marks

1. Calculate equivalent weight of Hydrogen peroxide and Oxalic acid.
2. Calculate equivalent weight of Potassium Permanganate and Iodine.
3. Complete and balance the equation: $KMnO_4 + H_2SO_4 \rightarrow$
4. Define Qualitative Analysis and Quantitative Analysis.
5. Differntiate between molar and normal solutions? What data is required to prepare these solutions.
6. Explain the importance of significant figures.
7. Give a list of methods of expressing concentration.
8. Give example for personal error and operative error.
9. Give the pH range of phenolphthalein and methyl orange indicators.

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10. How do you calculate stoichiometric end point in acid base titrations.
11. Name some indicators used in non-aqueous titrations.
12. Name the solvents used in non-aqueous titrations
13. Explain the uses of the following in non aqueous titrations a) perchloric acid b) acetic acid c) acetic anhydride d) crystal violet.

SESSION 2:

11/12/19 please ask the students write test by 3.00PM

5Marks

1. Explain the mechanism of action of indicators in Fajan's method.
2. Give the application of Gravimetric technique in the quantitative determination of barium as Barium sulphate.
3. Classify precipitation titration with examples. Explain Mohr's method in detail.
4. Discuss the principle and applications of Argentometric titrations with example.
5. Explain the principle and procedure involved in Volhards method and modified Volhards method.
6. Explain what is co-precipitation and post-precipitation with example.
7. Give the mechanism of action of adsorption indicators with suitable examples.
8. What is precipitation titration and give the principle involved in the assay of Sodium Chloride.
9. What is gravimetry? Explain the following terms a) Digestion b) ignition c) Ash treatment d) Inceneration
10. What is meant by Gravimetric analysis? Describe the techniques used for successful estimation of Barium.
11. Classify the various EDTA titrations and explain each one in detail.
12. Explain the principle and procedure involved in the estimation of Calcium Gluconate.
13. Explain the principle involved in the Complexometric titrations in detail and how will you estimate Magnesium Sulphate.

2Marks

1. Define co-precipitation and post-precipitation.
2. Difference between volhard's method and modified volhard's method.
3. What are Masking and Demasking Agents?
4. What are sequestering agents. Give examples.
5. What is back titration? Give example.
6. What is masking agent? Give example for masking by precipitation.
7. What is the difference between chelates and the complexes.
8. Write the structure of EDTA.
9. What are chelating agents. Give examples

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SESSION 3:

12/12/19 please ask the students write test by 3.00PM

10Marks

1. Discuss Iodometry and Iodimetry titrations briefly with examples.
2. Define oxidation and reduction. Give the applications of cerrimetry with suitable examples.
3. Define oxidising and reducing agents with suitable examples. Explain the principle involved in the iodometric titrations.
4. Define oxidation and reduction. Explain the principle involved in titration with potassium dichromate. Give its applications with suitable examples.
5. Classify redox titrations. Give the applications of cerimetry and bromatometry.
6. Give the principle and procedure involved in the estimation of ferrous sulfate and hydrogen peroxide.
7. Write the equation involed in the titration of iodine and sodium thiosulphate solutions. And explain the reactants and products.

2Marks

1. Give one example each for self indicator and internal indicator.
2. Give the difference between iodometry and iodimetry.
3. Give the formula to calculate the equivalents in redox titration.
4. What are self-indicators? Give examples.
5. What is back titration? Give example.
6. What is Cerimetry? Give its applications.
7. What is redox potential?
8. Give two example for redox indicators
9. Write the importance of Nernst equation.

SESSION 4:

13/12/19 please ask the students write test by 3.00PM

5Marks

1. Explain the conductometric titration curves for strong acid with weak base.
2. What is Ilkovic equation?derive an equation for it.
3. Write the construction and working of Glass electrode with advantages and disadvantages.
4. Enumerate the various types of electrodes in potentiometry. Give the working of Calomel electrode.

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5. Explain the construction and working of glass electrode. What are the advantages of glass electrode?
6. Explain the different steps involved in locating the end point in Potentiometric titrations.
7. Explain the titremetric curves obtained in conductometric titration a) strong acid Vs weak base b) strong base Vs strong acid.
8. Give the construction and working of DME.
9. Give the construction, working and applications of platinum electrode.
10. What are the reference and indicator electrodes used in potentiometric titrations. Explain construction and working of any one electrode.
11. What is polarographic curve? How it is ploted ? Mention different areas in the polarographic curves.
12. What is polarography? Explain the terms a) limiting current b) polarographic maxima c) diffussion current d) supporting electrolytes
13. Write the construction and working of conductivity cell.
14. Write the principle and applications of Polarographic analysis.
15. Write the principle, instrumentation and applications of conductometry.

2Marks

1. Define conductance and resistance.
2. Define Molar conductance and Specific conductance.
3. Give one example each for indicator electrode and reference electrode.
4. Name two compounds which can be estimated by conductometry.
5. What are Reference and Indicator electrode?
6. What is Null point potentiometry?
7. What is Specific conductance and Molar condutance?
8. What is standard hydrogen electrode?
9. Write the importance of Nernst equation.
10. What is conductivity cell.
11. Write the differences between conductometry and potentiometry.

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KRUPANIDHI COLLEGE OF PHARMACY, BANGALORE
QUESTION PREDICTABILITY ANALYSIS
Pharmaceutical Analysis-I

Teacher In-charge: Dr.P.D.Chaithanya Sudha

Report: Ph.Analysis- I, Jan 2020

Weightage	Number of given Question in set of 30 Q&A	Total Marks (out of 95)
Long Essay	All 3	30
Short Essay	8 questions	40
Short answers	8 questions	16
TOTAL QUESTIONS FROM SET OF 30 Q& A		86 marks out of 95

University Questions:2nd Jan 2020		Asked in IA Exam	Remarks	Marks (out of 95 marks)
LONG ESSAY				
1.	Errors	1st Sessional QP Additional Sessional QP Pre-final Examination QP	Given in 30 Q&A set	10
2.	Non-Aqueous titrations	1st Sessional QP Additional Sessional QP Pre-final Examination QP	Given in 30 Q&A set	10
3.	Redox titrations	II nd sessional QP Remedial classes Pre-final Examination QP	Given in 30 Q&A set	10
SHORT ESSAY				

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4.	Redox titrations	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	10
5.	Acid Base Titrations	1st Sessional QP Pre-final Examination QP	Given in 30 Q&A set	05
6.	Acid Base Titrations	1st Sessional QP Pre-final Examination QP	Given in 30 Q&A set	10 02
7.	Precipitation Titrations	Presessional, class tests and Work book	Not Given in 30 Q&A set	
8.	Complexometric titrations	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	05 02
9.	Gravimetric Analysis	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	05
10.	Potentiometry	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	05
11.	Conductometry	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	05
12.	Polarography	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	05
SHORT ANSWERS				
13.	Primary and secondary standards	I st sessional QP Pre-final Examination QP	Given in 30 Q&A set	05

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14.	Acid Base Indicators	I st sessional QP Pre-final Examination QP	Given in 30 Q&A set	02
15.	Equivalents weights calculation	I st and II nd presessional and work book		
16.	Non Aqueous titrations	I st sessional QP Pre-final Examination QP		
17.	Complexometric titrations	1I nd Sessional QP Pre-final Examination QP	Given in 30 Q&A set	05
18.	Redox titrations	1I nd Sessional QP Pre-final Examination QP	Given in 30 Q&A set	10
19.	Gravimetry	1I nd Sessional QP	Given in 30 Q&A set	02
20.	Electro analytical methods	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	05
21.	Conductometry	II nd sessional QP Pre-final Examination QP	Given in 30 Q&A set	02
TOTAL WEIGHTAGE from Q&A in set of 30 86 /95				

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I. Pre-Final QP Vs. University QP

KRUPANIDHI COLLEGE OF PHARMACY, BANGALORE
Pharmaceutical Analysis –I
I SEM B.PHARM, PRE-FINAL
EXAMINATION, DECEMBER 2019

MM: 75 marks **Time: 03:00 hrs**

Long Essay (Any two) **(10×2=20)**

Q1. Define and classify errors with examples and explain the methods to minimize the errors.
Q2. What are non aqueous titrations? Classify non aqueous solvents and explain the principle and procedure involved in the assay of sodium benzoate.
Q3. Define oxidizing and reducing agents with suitable examples and explain the different types of redox titrations with examples.

Short Essay (Any Seven) **07 x 05 = 35 marks**

Q4. What type of compounds is considered as primary and secondary standards? Give examples.
Q5. Explain the adsorption indicators with examples.
Q6. Classify and explain the various EDTA titrations.
Q7. Explain the construction and working of glass electrode.
Q8. What is gravimetry? Give the principle and procedure involved in the barium sulfate estimation.
Q9. Explain the principle involved in the estimation of strong acid with strong base with the help of titration curve.
Q10.Enumerate the various types of electrodes used in potentiometry and add a note on working of calomel electrode.
Q11.How do you prepare and standardize the 0.1N perchloric acid?
Q12. What is polarographic curve and mention the different areas in the polarographic curve.

Short Answer (Answer All) **02x 10 = 20 marks**
Q13. What are universal and mixed indicators?
Q14. What is standard hydrogen electrode?
Q15. Define neutralization curve and give its importance.
Q16. Define qualitative analysis and quantitative analysis.
Q17. Explain the uses of perchloric acid, acetic acid, acetic anhydride and crystal violet.
Q18. Give the P^H range of phenolphthalein and methyl orange.
Q19. Write the difference between conductometry and potentiometry.
Q20. Give the role of starch indicator in redox titrations.
Q21. What is specific and molar conductance?
Q22. What are masking and demasking agents?

Rajiv Gandhi University of Health Sciences, Karnataka
First Semester B. Pharm Degree Examination – 04-Jan-2020

Time: Three Hours Max. Marks: 75 Marks

PHARMACEUTICAL ANALYSIS - I
Q.P. CODE: 5002

Your answers should be specific to the questions asked.
Draw neat labeled diagrams wherever necessary.

LONG ESSAYS (Answer any Two) **2 x 10 = 20 Marks**

1. Define error; classify determinate error with suitable examples. Explain the terms 'accuracy' and 'precision'.
2. Write a note on solvents used in non-aqueous titrations. Explain the preparation and standardization of 0.1N perchloric acid.
3. Define oxidizing and reducing agents with a suitable example each. Discuss the principle of redox titrations. Explain standardization of 0.1N sodium thiosulphate solution.

SHORT ESSAYS (Answer any Seven) **7 x 5 = 35 Marks**

4. Define 'normal solution'. Explain preparation and standardization of 0.1N potassium permanganate solution (Mol. Wt: 158)
5. Explain the titration curve of strong acid versus strong base. How are these curves useful in titrimetric analysis?
6. Write a note on universal indicators and mixed indicators with examples and their uses.
7. Explain Mohr's method of determination of halides.
8. With a suitable example each, explain the terms 'masking', 'demasking', 'ligand' and 'chelate' in complexometric determinations.
9. Define gravimetry. Mention two compounds assayed by gravimetry. Explain the advantages and disadvantages of this technique.
10. Explain the construction and working of a glass membrane electrode.
11. Explain any two conductometric titration curves.
12. Define polarography and indicate its applications. Enumerate the Ilkovic equation.

SHORT ANSWERS **10 x 2 = 20 Marks**

13. With an example, define primary standard substance. Give its significance.
14. Mention two neutralization indicators, which work in acidic pH along with their pH interval respective colours.
15. Define equivalent weight of: 'base' and 'reducing agent' with an example each.
16. Illustrate effect of temperature in non-aqueous titrations.
17. Name four complexometric indicators.
18. Differentiate between 'iodometric' and 'iodimetric determinations'.
19. Short note on 'ignition' and 'peptization'.
20. How does starch act as an indicator in iodimetric titrations?
21. Differentiate between reference electrode and indicator electrode.
22. Define molar conductivity.

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KRUPANIDHI COLLEGE OF PHARMACY

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30 QUESTIONS AND ANSWERS FOR WEAK STUDENTS
Pharmaceutical Analysis-I

10 Marks

1. Define and classify errors with examples and list the methods for the minimization of errors.
2. Classify Acid-Base titrations and explain the theories of Indicators.
3. What are Non-Aqueous titrations and classify Non-Aqueous solvents? Write the Assay of Sodium Benzoate.
4. Define Oxidising agents and reducing agents with suitable examples and explain different types of redox titrations.
5. How do you prepare and standardize 0.1N perchloric acid and Calculate equivalent weight of a) H_2O_2 and b) Oxalic acid.

5 Marks

1. What types of compounds are considered as primary standards and give suitable examples.
2. Explain the principle involved in the estimation of mixture of strong acid with strong base with the help of titration curve.
3. Explain the adsorption indicators with examples.
4. Explain the principle involved in the Volhard's method and modified Volhard's method.
5. Classify and explain various EDTA titrations.
6. Why gravimetric estimation is preferred for certain compounds? Give the principle and procedure involved in Barium Sulfate estimation.
7. Explain the titrimetric curves obtained in conductometric titrations
   a) Strong acid with strong base
   b) Strong base with strong acid
8. Explain the construction, working and advantages of glass electrode.
9. What is polarographic curve? How it is plotted and mention different areas in the polarographic curve?
10. Enumerate the various types of electrodes in potentiometry and working of calomel electrode.

2Marks

1. Write the importance of Nernst equation.
2. What is specific and molar conductance?
3. What is standard hydrogen electrode?
4. Give the role of starch as Indicator in Redox titrations?
5. What is cerimetry and give its applications?
6. What are chelating agents?
7. What is masking agent?
8. Define neutralization curve and give its importance.
9. What are Universal or Mixed indicators?
10. Define quantitative and qualitative analysis.
11. Give the P^H range of phenolphthalein and methyl orange indicators.
12. Explain the uses of Perchloric acid, Acetic acid, Acetic anhydride and crystal violet.
13. What are conditions for process of digestion in gravimetry?

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14. What is the difference between chelates and complexes?
15. Write the difference between conductometry and potentiometry.

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SESSIONAL MARKS, I SEM B PHARM 2019-20

PHARMACEUTICAL ANALYSIS

Sl. No	STUDENT NAME	PH. ANALYSIS	
		I Sessional	II Sessional
		Theory (30)	Theory (30)
1	Muhammed Shahazain Sifath	12	21
2	Abhilash S. Reddy	25	29
3	Aditya Dholi	7	4
4	Akash	26	28
5	Arnold Joseph	20	21
6	Benson Babu Thomas	28	30
7	Bhavana L	25	24
8	Bhavya Shree Balasubramani	27	30
9	Bhoomika HV	20	29
10	Bhuvan R	14	21
11	Chaithra MS	24	26
12	Channa Basavana Gouda Patil	25	25
13	Charan	24	27
14	Chephar Samati	12	26
15	Deekshitha B	22	22
16	Deepika	27	30
17	Dilip Kumar GM	12	26
18	Durga Prasad V	18	28
19	Girinath M	14	25
20	Harini R	21	24
21	Harshini S	15	26
22	Harshita Maheshwari	21	22
23	HK Lalengzuala	AB	AB
24	Jagadish K	28	30
25	Chethan.H	3	AB
26	Jayashree R	29	29
27	Jayashree S	15	21
28	Jerome David	25	28
29	Jyothi Shruthi B K	15	21
30	Kavita Donniayavar	21	28
31	Kavya B	18	27
32	Kavya.HS	24	28
33	Keerthana S	29	29
34	Khaled Gamal Abdullah Mayas	AB	AB

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35	Kishore Kumar V	8	28
36	Krupa K	26	30
37	Lahari K	23	22
38	Lavanya M	14	26
39	Likhitha HR	15	21
40	Likitha M	20	28
41	M chandan	20	22
42	M.Abinash	13	24
43	M.Divyashree	28	30
44	M.Rakesh	25	24
45	M.Sanjay	18	24
46	Madan M	22	23
47	Makeshwar PE	30	30
48	Mallikarjun R	16	23
49	Manjunath R	14	21
50	Manjuprasad.M	23	27
51	Mithan Surya M	0	AB
52	Monica C	27	30
53	Monica S	28	29
54	Muhamed Hilal	22	27
55	Nandini M R	28	28
56	Nikhil E	22	26
57	Nikitha V	22	28
58	Nishu. BS	21	28
59	Nithya A	15	27
60	Nivya Anusha Reddy	22	30
61	Pavithra R.	21	27
62	Prajval M	17	22
63	Praveen Paul	16	21
64	Preethi AJ	30	29
65	Protik Kishore Choudhury	AB	AB
66	Purvika.K.S	25	29
67	Rafiya	26	29
68	Rahul Anganyan	15	29
69	Ranajit Sen	21	25
70	Ranjana devi sajjan	14	30
71	Ravi Kumar Sah	12	21
72	Rithiksha V	23	27
73	S.Fazila Parveen	19	28
74	S.Tejaswini Patil	23	21
75	Samker Kyizom	17	26
76	Sangeetha.J	AB	AB

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77	Santhip S	28	30
78	Saya Jyothi U	26	26
79	Shireesha KR	20	27
80	Shirisha R	28	30
81	Shreelekha L	19	21
82	Sibi M	25	29
83	Sridhar P	27	29
84	Sunil K	21	22
85	Sunil Kumar S	29	30
86	Sushma D	26	28
87	Sushma.N	25	30
88	Swetha K	18	24
89	Tejashwini J	29	30
90	Tejashwini K	29	30
91	Tenzin Nordon	22	22
92	Tenzin Woeser	18	28
93	Vaishnavi	18	28
94	Vanshika	25	25
95	Vasudeva M	26	30
96	Venu M	12	29
97	Vignesh S	22	26
98	Vinay Kumar N	29	30
99	Yashwanth.S	28	30
100	Yukitha.B	30	30

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REMEDIAL CLASSES 2018-19

Class: III Pharm D

Name of the students: Mohammad Asir, Siddarth M

SUBJECTS	Dates								Signature of teacher
Pharmacology II	22/04/2019	23/04/2019	24/04/2019	25/04/2019	26/04/2019	27/04/2019	29/04/2019	30/04/2019	
Pharmaceutical analysis	22/04/2019	23/04/2019	24/04/2019	25/04/2019	26/04/2019	27/04/2019	29/04/2019	30/04/2019	
Pharmacotherapeutics II	22/04/2019	23/04/2019	24/04/2019	25/04/2019	26/04/2019	27/04/2019	29/04/2019	30/04/2019	
Pharmaceutical jurisprudence	22/04/2019	23/04/2019	24/04/2019	25/04/2019	26/04/2019	27/04/2019	29/04/2019	30/04/2019	
Medicinal chemistry	22/04/2019	23/04/2019	24/04/2019	25/04/2019	26/04/2019	27/04/2019	29/04/2019	30/04/2019	
Pharma. Formulations	22/04/2019	23/04/2019	24/04/2019	25/04/2019	26/04/2019	27/04/2019	29/04/2019	30/04/2019	

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DETAILS OF REMEDIAL CLASSES

I PHARM D-2018-19

SL. NO	STUDENT NAME	HAP			PCEUTICS		MBC		POC			PIC		
		2/05/19	3/05/19	4/05/19	16/05/19	17/05/19	13/05/19	14/05/19	6/05/19	8/05/19	9/05/19	10/05/19	11/05/19	12/05/19
1.	Aishwarya Rakesh	P	P	P	P	P	P	P	P	P	P	P	P	P
2.	Anika Varsha Thamra S	A	A	A	A	A	A	A	A	A	A	A	A	A
3.	Divya M	P	P	P	P	P	P	P	P	P	P	P	P	P
4.	Lomash Timsina	P	P	P	P	P	P	P	P	P	P	P	P	P
5.	Moulidharan R	A	A	A	A	A	A	A	A	A	A	A	A	A
6.	Pawan Kumar K	A	A	A	A	A	P	P	P	P	P	P	P	P
7.	Sahana G	P	P	P	A	A	A	A	A	P	P	P	P	P
8.	Shiwangi Bhandari	P	P	P	P	P	P	P	P	P	P	P	P	P
9.	Sobin Thomas	A	A	A	A	A	A	A	A	A	A	A	A	A
10.	Rashmitha	P	P	P	P	P	P	P	P	P	P	P	P	P

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KRUPANIDHI COLLEGE OF PHARMACY, BANGALORE
REMEDIAL CLASSES, March 2019
FIRST SEMESTER B.PHARM, ATTENDANCE OF MENTEES

Subject: ~~Pharmaceutical~~ analysis /HAP/pharmaceutics/PIC. **Teacher In-Charge:**

Students' Name	08/03/19 (HAP)	09/03/19 (Analysis)	11/03/19 Pharmaceutics-I	12/03/19 (PIC)	13/03/19 HAP.
Sahana A	P	P	P	P	A
Santhala Chaithanya Prasad	A	P	A	P	P
Sasikirana S	P	A	P	A	P
Shiv Kumar M	P	P	P	P	P
Shivangi Chaudhary	P	P	P	P	P
Siddharth Bairaagi	P	P	A	A	P
Snigdha Senapati	P	P	P	P	A
Sobana S	P	A	P	P	P
Suchitra G	P	P	P	A	P
Suravi Dutta	P	P	P	P	P
Swetha M	P	P	A	P	P
Tamil Selvan	A	A	P	P	A
Tathagat Singh	P	P	A	P	P
Uday Kiran	P	P	P	P	P
Velavan P	A	P	A	A	P
Venhatachalam M	P	P	P	P	A
Venu G	A	A	A	P	P
Vijaylakshmi M	P	P	P	P	P

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13/03/19.

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FIRST SEMESTER B.PHARM, ATTENDANCE OF MENTEES

Subject:PIC

Teacher In-Charge:
Arnika Das

	Students' Name	8/3/19	9/3/19	11/3/19	12/3/19	13/3/19
1.	Ajith	✓	✓	✗	✓	✓
2.	Abhilash	✓	✓	✓	✓	✓
3.	Aleena	✓	✓	✗	✓	✓
4.	Azzam	✗	✓	✓	✓	✓
5.	Ashwini R	✓	✓	✓	✓	✓
6.	Kavya nikitha	✗	✓	✓	✓	✓
7.	Anisha	✗	✗	✓	✓	✓
8.	Anusha	✗	✓	✓	✓	✓
9.	Ashiwini Ramaswamy	✓	✓	✓	✓	✓
10	Chandana	✓	✓	✓	✓	✗
11	Babul	✓	✓	✓	✓	✓
12	Binuka	✓	✓	✓	✓	✓
13	Chaitanya	✓	✓	✓	✓	✗
14	Yaseen	✓	✓	✓	✓	✓
15	Angelin	✓	✓	✓	✓	✓
16	Rashmi	✓	✓	✗	✓	✓
17	Aeksa	✓	✓	✓	✓	✓

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